or ati-» cubic metres 0()a,so one obtains from 1 kilogramme C, 22*;i
0 - = 1'H() cubic moire, C015 (under normal conditions 0° and
7(K) millimetres) . . . (4).
MHO when combustion occurs by moans of tho theoretically requisite amount of air, that N=H-7<>'2 (H)o-. ;l"7(J2x l-HO-T'OO cubic metres N. Thus tho volume of tho gases of combustion of 1 kilogramme C \vilh air under normal conditions is
1 -HO cubic metro (•().,+ 7-00 cubic metres N ^H'BG cubic
mot/res «, * (f>).
(It) 00 FORMATION. Aiu-CUs.1
If one considers tho formation of CO according to tho equation C [O CO, then H kilogmmmo.H C with UJkilo^rammosO yi(*ltl *2K kilo^ratumes 0.). On division by tlu^ rcmpoctivo woightB
of 1 cubic metre. O and (U), one, obtains
O ami
••- 11*15 cubic ic moires CO. Thus 11*15
cubic, mrtroH O yields 22w»i cubic metres CO, and from this it follows that I cubic metro 0 yields "2 cubic, metres CO, or that for producing I cubic motro CO, -h cubic metro 0 is rtujiurod. . . ((>)•
For tho manufacture of tho gas, air Is again employed, thus there will be yielded from 21 cubic metres O, :!Xk21 --1*2 cubic mntroH CO and 7*J cubic metres N unchanged. Thus 100 cubic motren uir will yi(dd 1421 cubic mtstros gas. This will
, n -4- 42X100 ,Mrr , /1n ^tlxlOO
havit the compoHittou -•• U4"7 volinnoH (A), and fc>1
I *j I I
*j A,
— fia'H volumes N . . . (7).
In thin gan, the nitrogen contont in defined according to the
formula N - :)'7ri*2
CO '
(B), as in tho production of
1 cubic metro CO only A cubic nu»tn* O IB nocoHnary.
itif u»n ttW«'»i!', j
hlntor hiiH iMitip
. Tin- Ut»n»«ui i*xj»rf!HNion in 4* Ijuft"pw," uml t-hi*, lim* tn Ui«* Fortnnla nuinlH-rrd 7, in so awful l.lml ilm int lit mil tnuihiiditju of, iu»<l <U?tUiitio» of this («cnuun
wurti.